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**Electrochemical determination of traces lead ions using a new nanocomposite of
polypyrrole/carbon nanofibers**

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Abstract

In this paper, a new nanocomposite of polypyrrole (PPy) and carbon nanofibers (CNFs) modified carbon paste electrode (CPE) has been reported for the determination of traces lead ions (Pb^{2+}). The obtained nanocomposite was fabricated by combining the unique advantages of PPy and CNFs using a very simple approach, which consists on modifying the CPE by the functionalized CNFs and then by the PPy film using galvanostatic mode. Several techniques were used to investigate the functionalized CNFs and the PPy/CNFs nanocomposite including Fourier transform infrared spectroscopy (FTIR) and X-ray diffraction (XRD) for surface layers of CNFs, cyclic voltammetry (CV) and electrochemical impedance spectroscopy (EIS) for the electrical proprieties of the PPy/CNFs nanocomposite. The surface morphologies were examined by field emission gun scanning electron microscopy (FEG-SEM). The square wave anodic stripping voltammetry (SWASV) was used to investigate the analytical performances of the designed electrode PPy/CNFs/CPE. Different parameters that affect the stripping analysis of Pb^{2+} including supporting electrolyte, deposition potential, and deposition time were investigated. Under the optimum experimental conditions, a good linearity between the stripping peak currents and the concentration of Pb^{2+} was obtained in the range of concentration from 0.2 to 130 $\mu\text{g L}^{-1}$ Pb^{2+} . The detection limit was estimated to be 0.05 $\mu\text{g L}^{-1}$ Pb^{2+} . Finally, the proposed method has been successfully applied for the determination of Pb^{2+} in real samples of tap water with satisfactory results.

Keywords: Polypyrrole; Carbon nanofibers; Carbon paste electrode; Lead ions; Square wave voltammetric

Introduction

Nanocomposites are a class of materials, formed by combining multiple phases at nonmetric scale, with the aim to produce new materials with interesting proprieties. Over the last decades nanocomposites based on carbon nanomaterials (CNMs) including carbon nanotubes (CNTs), graphene (GR) and carbon nanofibers (CNFs) combined with conducting polymers (CPs) such as polypyrrole (PPy), polyaniline (PANi), polythiophene (PTh), and poly(3,4-ethylenedioxythiophene) (PEDOT) have opened a new era in the field of nanoscience and nanotechnology [1]. More recently, CNMs/CPs nanocomposites have received considerable attention as sensing materials due to their extraordinary proprieties mainly including excellent electrical conductivities, large active surface area and stability [2]. Monitoring and control of environmental pollution are fields that these nanocomposites have been extensively used as sensors to analyze a variety of pollutants such as heavy metals ions (Pb^{2+} , Cd^{2+} , Hg^{2+}) and pesticides [3, 4]. In addition, to overcoming the new challenges posed by different samples matrices. Traces detection of heavy metals ions in water is necessary due to high toxicity and serious risks posed by these pollutants for environmental and human health. Among of them, lead ions (Pb^{2+}) is one of the most toxic heavy metals ions even at a low concentration exposition [5, 6]. Considerable efforts have been paid to develop a simple and a highly sensitive analytical methods for the analysis of lead ions and many other heavy metals in water [7, 8]. Electrochemical methods including square wave anodic stripping voltammetry (SWASV) using modified electrodes provide more attractive feature as suitable tools in the sensing applications, due to their great potentials which are easy to operate, low cost, high sensitivity, and selectivity [9]. Several research works have reported the use of electrochemical methods and CNMs/CPs nanocomposite modified electrodes as sensing platform for heavy metals ions. Graphene film electrode modified with PANi and bismuth nanoparticles was used to detect traces of Cd^{2+} and Pb^{2+} using SWASV [10]. A nanocomposite of multi-walled carbon nanotubes/poly(1,5-

diaminonaphthalene) (MWCNT/P-1,5-DAN) modified platinum electrode was reported for the analysis of Pb^{2+} and Cd^{2+} using SWASV. The modified electrode showed a low detection limit estimated around of 3.2 and 2.1 $\mu\text{g L}^{-1}$ respectively for Cd^{2+} and Pb^{2+} [11]. Another nanocomposite of P-1,5-DAN and reduced graphene oxide (rGO) modified platinum patterned electrode was fabricated and used for Pb^{2+} determination. The detection limit of Pb^{2+} using the P-1,5-DAN/rGO/Pt was equal to 0.2 $\mu\text{g L}^{-1}$ [12]. PPy and carbonaceous nanospheres (CNSs) modified screen-printed electrode (SPE) was also fabricated for a selective detection of Hg^{2+} and Pb^{2+} using SWASV [13]. PPy and reduced graphene oxide (rGO) modified glassy carbon electrode (GCE) was reported as a new nanocomposite electrode for a high selective detection of Hg^{2+} [14]. Cysteine functionalized graphene oxide (sGO) and polypyrrole modified screen printed electrode (SPE) was reported and applied for the determination of Pb^{2+} using differential pulse anodic stripping voltammetry (DPASV) [15]. GCE modified with MWCNTs functionalized 2-(5-bromo-2-pyridylazo)-5-diethylaminophenol (MWCNTs/5-Br-PADAP) was prepared to fabricate a new electrochemical sensor for Pb^{2+} determination using DPASV [16].

Carbon paste electrodes (CPEs), have been widely used in a variety of electrochemical applications due to their interesting properties such as a low background current, low cost, a simple preparation and renewal of the surface [17]. However, the use of these electrodes as sensors requires an improvement of their properties. Chemically modified carbon paste electrodes (CMCPEs) have been introduced as a suitable approach to improve their analytical performances mainly including the sensitivity, selectivity, stability, and detection limit [18–21]. The CMCPEs consist of the modification of the surface of CPE by various materials such as metals nanoparticles (MNPs), CNMs, and CPs [22–24]. The combination of these materials was also applied to combine their fantastic proprieties. The resulting modified electrodes could be used for the analysis of several analytes including heavy metals ions, pesticides, and many other compounds [25–27].

Vapor grow carbon nanofibers (VGCNFs) or simply (CNFs) are a class of carbon nanomaterials with a cylindrical nanostructure of graphene layers arranged as stacked cones, cups or plates [28–30]. CNFs are recognized by their remarkable properties including excellent electrical and thermal conductivities, high chemical and mechanical stability, good specific surface area and a possibility of mass production. These properties make CNFs a very promising nanomaterial in various application area [31–34]. However, few studies have reported their application in electrochemical sensors.

In the present work a new nanocomposite of PPy/CNFs modified CPE was reported. A very simple approach was used to combine the unique properties of PPy and CNFs, consisting of the dispersion of the functionalized CNFs at the CPE surface then followed by the electropolymerization of PPy using galvanostatic mode. The PPy was used because of its remarkable properties including good electrical conductivity, high stability, and a facile synthesis. Moreover, PPy is a promising conjugated conducting polymer for the development of new electrochemical sensors, particularly for heavy metal ions. To our best knowledge, no work has been devoted to the CNFs and PPy as sensing materials for the detection of heavy metals ions. Different techniques were used to investigate the structural and electrical proprieties of the functionalized CNFs and PPy/CNFs nanocomposite including Fourier transform infrared (FTIR), X-ray diffraction (XDR), cyclic voltammetry (CV), electrochemical impedance spectroscopy (EIS), and field emission gun scanning electron microscopy (FEG-SEM). The analytical performances of PPy/CNFs/CPE for the determination of Pb^{2+} were studied using SWASV, several parameters affecting the stripping analysis were investigated. Finally, the application of the proposed method for the analysis of Pb^{2+} in real samples was tested with satisfactory results.

Materials and methods

Chemicals and reagents

Pyrrole Py (99%) was obtained from ACROS Organics and purified by distillation prior to use. Vapor grow carbon nanofibers (CNFs) with diameter of 100 nm and length range from 20 to 200 μm , graphite powder (particle size less than 20 μm), paraffin oil, lithium perchlorate LiClO_4 (purum p.a. 98%), lead (II) nitrate $\text{Pb}(\text{NO}_3)_2$ (ACS reagent, $\geq 99\%$) and potassium ferrocyanide-ferricyanide $\text{K}_3\text{Fe}(\text{CN})_6/\text{K}_4\text{Fe}(\text{CN})_6 \cdot 3\text{H}_2\text{O}$ (ACS reagent, $\geq 99\%$) were obtained from Sigma Aldrich. Sulfuric acid H_2SO_4 (98%) was purchased from Fluka, potassium chloride KCl (99%) and nitric acid HNO_3 (68%) from AnalaR NORMAPU. The acetate buffer solution was prepared by mixing an appropriate amounts of acetic acid sodium CH_3COONa (anhydrous ACROS Organics) and acetic acid CH_3COOH (Sigma Aldrich). All other chemicals and reagents used in this study were of analytical grade.

Materials and apparatus

The electrochemical measurements were performed in a conventional three-electrodes cell containing the bare or the modified CPE as working electrode, platinum grid with a surface area of 4 cm^2 as counter electrode and saturated calomel electrode (SCE) as reference electrode (Ref. 421, Radiometer Analytical). Voltammetry measurements including cyclic voltammetry (CV) and square wave voltammetry (SWV), were performed with an EcoChemie Autolab PGSTAT100 Potentiostat/Galvanostat under the control of GPES software, version 4.9. Electrochemical impedance measurements were performed using a potentiostat Sotalem and a frequency response analyzer (FRA, Solartron 1254) controlled by a software developed in the laboratory LISE. Fourier transform infrared spectroscopy (FTIR) were recorded with KBr pellets using Bruker IFS 66/S FTIR spectrometer. The X-ray diffraction (XRD) analysis were performed using a Panalytical Empyrean diffractometer equipped with a $\text{Cu-K}\alpha$ as radiation

source ($K\alpha = 1.54 \text{ \AA}$). The surface morphology of the modified electrodes was analyzed by field emission gun scanning electron microscope (FEG-SEM, Zeiss-Supra 55).

Functionalization of CNFs

The acid treatment is a method widely used to oxidize the CNFs or CNTs by creating carboxyl (-COOH), carbonyl (-CO), and hydroxyl (-COH) groups on their surface. In addition, this treatment could eliminate the impurities from the production process principally metal catalysts [35–37]. For that an amount of CNFs (1 g) was added to a mixture solution of concentrated nitric and sulfuric acid (3:1 volume ratio) and dispersed under ultrasound during 10 min. Then, the mixture solution was refluxed at 80 °C under magnetic stirring during 6 h. The residue was then filtered using a membrane (Millipore polycarbonate, 0.22 μm in diameter) and rinsed with bidistilled water until a neutral pH. Finally, the obtained CNFs were dried at 60 °C during 8 hours.

Preparation of modified electrodes

Carbon paste electrode was prepared by mixing the graphite powder (1 g) and the paraffin oil (300 μL) in a mortar until a homogeneous paste was obtained [38]. A portion of resulting paste was carefully packed into electrode cavity (Teflon tube, 3 mm in diameter), and then polished on a clean paper until a smooth surface was obtained. The electrical contact was established using a stainless steel rod with a diameter of 2 mm. CNFs modified CPE was prepared by dropping 10 μL of a suspension of the functionalized CNFs onto the CPE surface, which is prepared by dispersing 2 mg of the treated CNFs in 1 ml of the bidistilled water under ultrasonication. The obtained modified electrode was then dried at room temperature and designed CNFs/CPE. The PPy film modified bare CPE and CNFs/CPE was synthesized using galvanostatic mode in a solution of 0.1 M pyrrole and 0.5 M LiClO_4 at an applied constant current of 0.2 mA cm^{-2} during 100 seconds. The modified electrodes were denoted PPy/CPE

and PPy/CNFs/CPE thoroughly washed with bidistilled water and dried under a gentle stream of nitrogen gas.

Measurement procedure

Cyclic voltammetry (CV) measurements were performed at a potential scan rate of 50 mV s^{-1} in 0.25 M LiClO_4 , and in 10 mM of $[\text{Fe}(\text{CN})_6]^{3-/4-}$ solution containing 0.5 M KCl respectively at a potential range from $+0.3$ to -0.8 V/SCE and $+0.6$ to -0.4 V/SCE . Electrochemical impedance measurements were carried out in 10 mM of $[\text{Fe}(\text{CN})_6]^{3-/4-}$ solution containing 0.5 M KCl at equilibrium potential in the frequency ranges from 10 mHz to 63 kHz and an amplitude of 10 mV . SWASV measurements were performed in 0.1 M acetate buffer solution at $\text{pH } 4.5$ containing a predetermined concentration of targets metal ions. A deposition potential of -1.2 V/SCE was applied to the working electrode PPy/CNFs/CPE during 600 seconds and under magnetic stirring. After a period of stabilization of about 15 seconds, the square wave voltammograms (SWVs) were recorded by applying simultaneously a frequency of 50 Hz , step potential of 5 mV , amplitude of 50 mV , and a scan positive potential range from -0.9 to -0.3 V/SCE . To remove the possible residual metal ions on the electrode surface a potential of 0.3 V/SCE was applied to the working electrode during 15 seconds under magnetic stirring prior to the next measurement. All the experiments were carried out in a constant temperature ($25.0 \pm 0.1 \text{ }^\circ\text{C}$) under nitrogen gas environment.

Results and discussion

FTIR and XDR characterization of the functionalized CNFs

FTIR spectroscopy has been widely used as a suitable investigation to identify the surface functional groups. Fig. 1a shows the FTIR spectra of CNFs before and after acid treatment. It could be seen that the treated CNFs demonstrate new bands compared to untreated CNFs which

leading to new functional groups on their surface. The band appeared at 3430 cm^{-1} indicates the presence of the hydroxyl groups ($-\text{OH}$), while the adsorption bands at 2864 and 2927 cm^{-1} are respectively assigned to symmetric and asymmetric methylene stretching bands. The bands located at 2336 and 2361 cm^{-1} are attributed to hydrogen bonding due to carboxylic groups ($-\text{COOH}$) which are formed during the oxidation of $-\text{OH}$ groups. The formation of $-\text{COOH}$ groups on the treated CNFs surface could be confirmed by the appearance of a new band at 1741 cm^{-1} , in addition to the band at 1113 cm^{-1} which are respectively attributed to stretching vibrations of carbonyl groups $\text{C}=\text{O}$ and $\text{C}-\text{O}$ presented in $-\text{COOH}$ groups [39, 40]. The adsorption band at 1635 cm^{-1} is assigned to the conjugated $\text{C}=\text{C}$ stretching related to CNFs structure. The obtained results shows clearly that the $-\text{COOH}$ groups were successively created on the treated CNFs surface, which could be leads to a good dispersion and purification of CNFs [41]. XRD analysis were also used to control the crystalline structure of CNFs before and after acid treatment. The XRD patterns for the CNFs is shown in Fig. 1b, it could be seen that the CNFs shows almost a same first order diffraction peak located at $2\theta = 26^\circ$ which is attributed to the hexagonal structure of CNFs with a Miller indices of (002) and an experimental spacing d of 3.37 \AA according to Bragg's equation [42]. The difference in intensity of peaks could be assigned to the different adsorption coefficients of the functional groups at the CNFs surface, in addition to the graphitization degree after the acid treatment as discussed in the literature [43]. The obtained results affirm that the treated CNFs have been no huge change in their crystalline structure.

Fig. 1 a) FTIR spectrum and b) XRD patterns of CNFs before and after acid treatment

Electropolymerization of PPy and cyclic voltammetry (CV) characterization of the modified electrodes

Electrochemical polymerization of PPy on bare CPE and CNFs/CPE surface was performed using galvanostatic mode. Fig. 2a displays the chronopotentiometric curves obtained during the deposition of PPy. The potential decreases during the first seconds then stabilize at 0.6 and 0.5 V/SCE respectively for CPE and CNFs/CPE, which corresponds to the electropolymerization potential of PPy. The difference in the potential between the two electrodes could be attributed to the good electrical properties of the PPy on CNFs. It is supposed that the functional group of CNFs promoted charger transfer during the electropolymerization leading to reduction of the deposition potential. These results are in agreement with those reported by Shi and Zhitomirsky [44]. To confirm the successful synthesis of the PPy film, the modified electrodes were characterized in 0.25 M LiClO₄ using cyclic voltammetry (CV). Fig. 2b shows the cyclic voltammograms recorded at different modified electrodes at a scan rate of 50 mV s⁻¹. It could be seen that the bare CPE and CNFs/CPE are characterized with a very background current, while PPy film modified CPE and CNFs/CPE have shown two typical redox systems indicating that the PPy film was successfully electropolymerized on the both surfaces. However, in the presence of CNFs the anodic and cathodic peaks currents of the PPy were increased with a well-defined peaks. In addition, the peak potential separation (ΔE_p) was decreased, revealing the role of the CNFs to enhance the electronic properties of PPy film. These results are similar with those reported in the literature concerning the polymerization of PPy on CNTs at a constant current [45]. In order to complement our investigation, the modified electrodes were then characterized in 10 mM [Fe(CN)₆]^{3-/4-} solution containing 0.5 M KCl. Fig. 2c shows the cyclic voltammograms recorded at different modified electrodes at a scan rate of 50 mV s⁻¹. The bare CPE was characterized by a quasi-reversible redox reaction with a ΔE_p of about 293 mV, indicating a slow electron transfer at the CPE surface. On the CPE modified with CNFs the charge transfer rate was improved leading to a ΔE_p of about 110 mV, with an increase of current of about 150% compared to the bare CPE. This behavior could be attributed to the presence of

the functionalized CNFs by providing a much active surface and an easier electron transfer [46]. PPy film modified bare CPE and CNFs/CPE have shown a significant shift of the voltammetric wave with an increase of peak current. The ΔE_p observed at PPy/CPE and PPy/CNFs/CPE was respectively equal to 85 and 68 mV, revealing that the PPy and PPy/CNFs films has been assembled onto the CPE and enhanced the electron transfer at the electrode surface. PPy/CNFs modified CPE provides a high current around 270%, more than obtained at bare CPE, with a good faradaic redox behavior. This behavior could be attributed to good electrical and surface properties of the formed PPy/CNFs nanocomposite as reported in previous studies by combining CNMs and CPs such as PPy/CNSs [13], PPy/sGO [15] and P1,5-DAN/MWCNT [11].

Fig. 2 a) Chronopotentiometric curves of PPy on bare CPE (cyan line) and CNFs/CPE (green line) obtained at an applied constant current of 0.2 mA cm^{-2} from aqueous solution containing 0.1 M pyrrole and 0.5 M LiClO_4 . Typical cyclic voltammograms of the fabricated electrodes: bare CPE (red line), CNFs/CPE (blue line), PPy/CPE (cyan line) and PPy/CNFs/CPE (green line) recorded at scan rate of 50 mV s^{-1} **b)** in 0.25 M LiClO_4 **c)** in 10 mM $[\text{Fe}(\text{CN})_6]^{3-/4-}$ solution containing 0.5 M KCl

Electrochemical impedance spectroscopy (EIS) characterization

EIS technique is a powerful technique to investigate the interface properties of the modified electrodes. The impedance measurements were recorded out in 10 mM $[\text{Fe}(\text{CN})_6]^{3-/4-}$ solution containing 0.5 M KCl at the equilibrium potential in the range of frequency from 10 mHz to 63 KHz. Fig. 3 shows the typical Nyquist plots achieved at different modified electrodes, and the estimated values of the charge transfer resistance (R_{ct}) and the double layer capacitance (C_{dl}) which are obtained by fitting the impedance data using Randles equivalent circuit. At bare CPE and CNFs/CPE it could be seen that the R_{ct} decreases respectively from 32.0 to $5.9 \text{ } \Omega \text{ cm}^2$

indicating a poor electron transfer reaction of $[\text{Fe}(\text{CN})_6]^{3-/4-}$ at the bare CPE, while the C_{dl} increases respectively from 6 to 255 $\mu\text{F cm}^{-2}$. Indeed, the presence of the non-conductive paraffin oil at the electrode surface decrease the active surface area and then the electron transfer [47]. However, the electrical properties of CPE were improved after its modification by the functionalized CNFs, inducing their role to promote the electron transfer as reported previously using CNFs and other CNMs [11, 12, 46, 48]. In the case of PPy film modified CPE and CNFs/CPE, the R_{ct} value decreased even more compared to these electrodes before their modification. These results showed that the PPy was successfully electropolymerized on the modified electrodes and could promote the electron transfer reaction of $[\text{Fe}(\text{CN})_6]^{3-/4-}$ due to its interesting electronic proprieties [13]. In particular at the PPy/CNFs/CPE, the R_{ct} has been further reduced and reach a value of 0.3 $\Omega \text{ cm}^2$ with increase of the C_{dl} , indicating the improvement of conductivity of the film. This improvement is due to the contribution of the large surface area offered by CNFs during electropolymerization as reported by others authors using nanocomposite such as PPy/sGO [15], PPy/MWCNTs [49]. Based on the R_{ct} and C_{dl} values of the fabricated electrodes, we can concluded that the electron transfer rates at the PPy/CNFs/CPE was the fastest, indicating the role of the synergic effects of PPy and CNFs to promote electron transfer reaction of $[\text{Fe}(\text{CN})_6]^{3-/4-}$. The EIS measurements are in good agreement with those obtained with CV.

Fig. 3 Nyquist plots of bare CPE (red dots), CNFs/CPE (blue dots), PPy/CPE (cyan dots), and PPy/CNFs/CPE (green dots) in 10 mM $[\text{Fe}(\text{CN})_6]^{3-/4-}$ and 0.5 M KCl in frequency range from 63 kHz to 10 mHz at an applied potential of +0.2 V/SCE. Inset is the simulation results of the EIS spectra

SEM characterization of the prepared electrodes

The surface morphologies of different modified electrodes was characterized by FEG-SEM as shown in Fig. 4. The bare CPE (Fig. 4a) was characterized by a surface formed by irregularly shaped flakes of graphite isolated from each other with a micrometer-size [50]. CNFs modified CPE (Fig. 4b) displays a uniform dispersion of CNFs on the electrode surface, and forming a spaghetti-like porous reticular formation which can offer a much active surface than in the case of bare CPE [48, 51]. CPE coated with PPy film (Fig. 4c) showed a cauliflower-like morphology with a micrometric grain size, while in the case of PPy film modified CNFs/CPE (Fig. 4d) it could be seen that the PPy was uniformly deposited on the CNFs surface revealing a finer and more compact structure.

Fig. 4 SEM images of a) bare CPE, b) CNFs/CPE, c) PPy/CPE and d) PPy/CNFs/CPE

Anodic stripping behavior for Pb^{2+} determination

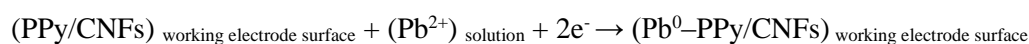
The analytical performances of the fabricated electrodes were tested using SWASV for the determination of $100 \mu\text{g L}^{-1} \text{Pb}^{2+}$ in 0.1 M acetate buffer solution at pH 4.5. Fig. 5 showed the square wave voltammograms (SWVs) recorded at different modified electrodes after a preconcentration time of 180 seconds, and a deposition potential of -1.2 V/SCE.

Fig. 5 SWVs of $100 \mu\text{g L}^{-1} \text{Pb}^{2+}$ at different modified electrodes. Deposition potential -1.2 V, deposition time 180s, frequency 50 Hz, amplitude 50 mV, step potential 5 mV and supporting electrolyte 0.1 M acetate buffer (pH 4.5)

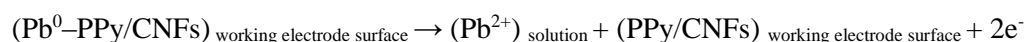
An anodic stripping peak was obtained at different modified electrodes, corresponds to the reoxidation of the reduced lead metals (Pb) to lead ions (Pb^{2+}). It could be seen that the best response was obtained at PPy/CNFs/CPE. The peak current of Pb^{2+} was improved seven times more than in the case of bare CPE. The enhanced stripping currents of Pb^{2+} at PPy/CNFs/CPE is due to a larger amount of Pb^{2+} reduced at its surface. Indeed, the PPy/CNFs nanocomposite modified CPE can provide a high active surface area and good electrical conductivity, which

allows an easier electron transfer between the electrode and the nanocomposite film. In addition, to the ability of PPy film to chelate the metal ions via amino groups along its macromolecular chain as in the case of some others conjugated conducting polymers such as 1,8-diaminonaphthalene, 1,5-diaminonaphthalene and polyaniline [11, 52–54]. These proprieties make the PPy/CNFs nanocomposite as a perfect sensing material for traces Pb²⁺ detection. At bare CPE and modified with CNFs we note a slight shift of peak potential of Pb²⁺ which could be attributed to the adsorption phenomena on their surfaces [46]. The overall mechanism of accumulation and anodic stripping of Pb²⁺ at PPy/CNFs/CPE using SWASV could be provided as the fallowing:

Accumulation step:



Anodic stripping step:



Optimization of the experimental conditions

Effect of supporting electrolytes and SWASV conditions

It is well known that during the stripping analysis of heavy metal ions using SWASV, the peak response is directly proportional to their concentration. Therefore, the optimization of the experimental conditions including supporting electrolytes and SWASV conditions is very important to improve the sensitivity of the method. The effect of supporting electrolytes was studied in different solutions: 0.1 M acetate buffer (CH₃COOH/CH₃COONa), 0.1 M phosphate buffer (NaH₂PO₄/Na₂HPO₄), and 0.1 M hydrochloric acid (HCl). The anodic peak currents for the detection of 100 µg L⁻¹ Pb²⁺ using PPy/CNFs/CPE showed that a high current response was obtained in 0.1 M acetate buffer at pH 4.5. Therefore, it was selected as an appropriate supporting electrolyte for the next experiments. Previous studies showed that the acetate buffer

was a commonly supporting electrolyte used in anodic stripping voltammetry of Pb^{2+} [11, 55]. Then, the SWASV parameters including frequency, step potential, and amplitude were optimized using PPy/CNFs/CPE in 0.1 M acetate buffer at pH 4.5 as supporting electrolyte. The analysis of $100 \mu\text{g L}^{-1} \text{Pb}^{2+}$ revealed that a high stripping peak current was obtained using simultaneously a frequency of 50 Hz, a step potential of 5 mV and an amplitude of 50 mV.

Effect of deposition potential

The deposition potential is an important parameter in stripping analysis, which can strongly affect the sensitivity of the method. The effect of this parameter on the stripping signal of Pb^{2+} using the modified electrode PPy/CNFs/CPE was investigated in the range of potential from -0.8 to -1.6 V/SCE. Fig. 6 shows a plot of stripping peak currents of $150 \mu\text{g L}^{-1} \text{Pb}^{2+}$ versus the deposition potential. These results were recorded in 0.1 M acetate buffer pH 4.5 at a deposition time of 180 seconds. The stripping peak currents of Pb^{2+} were gradually increased when the deposition potential shifted from -0.8 to -1.3 V/SCE, a high stripping peak current was obtained at -1.3 V/SCE. However, when the deposition potential become more negative the stripping peak currents of Pb^{2+} were greatly decreased. This behavior is due to the evolution of hydrogen at the electrode surface, which could influence the active surface area leading to a decrease of amount Pb^{2+} reduced on the electrode surface [51]. Consequently, a potential of -1.2 V/SCE was selected as an optimum deposition potential for the following experiments.

Fig. 6 Effect of deposition potential on the stripping peak current of $150 \mu\text{g L}^{-1} \text{Pb}^{2+}$ in 0.1 M acetate buffer (pH 4.5) at PPy/CNFs/CPE. The SWVs were recorded at frequency of 50 Hz, amplitude of 50 mV and step potential of 5 mV

Effect of the preconcentration time

The influence of the preconcentration time was investigated from 2 to 24 min in 0.1 M acetate buffer solution pH 4.5 containing $150 \mu\text{g L}^{-1} \text{Pb}^{2+}$ at a deposition potential of -1.2 V/SCE. Fig.

7 shows a plot of peak currents of Pb^{2+} versus the preconcentration time. The stripping peak currents of Pb^{2+} are almost linear with the increase of the preconcentration time between 2 to 18 min, which is attributed to the increase of amount Pb^{2+} reduced on the electrode surface. These results clearly demonstrate that the PPy/CNFs nanocomposite is characterized with a high capacity to the accumulation the reduced metal ions due to its large active surface area. At a preconcentration time greater than 18 min, the electrode surface begins to saturate. It is known that a long preconcentration time can improve the sensitivity of the method and also decrease the detection limit, because of rapid saturation of the electrode surface, then a preconcentration time of 10 min is therefore selected as a compromise between a high signal and a suitable time for the next experiments.

Fig. 7 Effect of deposition time on the stripping peak current of $150 \mu g L^{-1} Pb^{2+}$ in 0.1 M acetate buffer (pH 4.5) using PPy/CNFs/CPE at deposition potential of -1.2. Other experimental conditions are mentioned in Fig. 7

Analytical characteristics

Under the optimized experimental conditions described above, the calibration curves for determination of Pb^{2+} at different concentration range from 0.2 to $130 \mu g L^{-1}$ using PPy/CNFs/CPE are illustrated in Fig. 8.

Fig. 8 Calibration curve for Pb^{2+} determination, from down to up: 0.2, 0.5, 1, 5, 15, 20, 30, 40, 60, 80, 100, and $130 \mu g L^{-1} Pb^{2+}$ in 0.1 M acetate buffer (pH 4.5) at PPy/CNFs/CPE under the optimized experimental conditions. Insert the calibration plot of oxidation peak current versus different concentrations of Pb^{2+}

It could be seen that the stripping peak currents (i_p) increase linearly with the concentration of Pb^{2+} (C), the calculated correlation equation is: $i_p = 0.289C - 0.0556$ ($R^2 = 0.9904$). The limit of detection (LOD) is calculated using the following equation: $LOD = 3\sigma/b$, where σ is the

standard deviation of the response, according to the IUPAC recommendations σ is the mean value for 10 voltammograms of the blank, and b is the slope of the calibration curve [56]. Therefore, the LOD of Pb^{2+} determination is found to be $0.05 \mu\text{g L}^{-1}$. The analytical performances of the proposed method are compared with other methods previously reported in the literature for Pb^{2+} determination are given in Table 1.

Table 1 Comparison of the analytical performance of PPy/CNFs/CPE with some other modified electrode for Pb^{2+} determination in aqueous media (detection limit and linear range are given in $\mu\text{g L}^{-1}$ except * in nM and \$ in μM)

The repeatability of the proposed method using PPy/CNFs/CPE under the optimized experimental conditions was evaluated by assaying $50 \mu\text{g L}^{-1} \text{Pb}^{2+}$. After five successive measurements using the same modified electrode, the calculated relative standard deviation (RSD) was around 2.3%. On the other hand, the reproducibility of the proposed method was also tested, a series of five electrodes were prepared and then applied for the detection of $50 \mu\text{g L}^{-1} \text{Pb}^{2+}$. The RSD of the measurements for the prepared five electrodes was equal to 3.1%, which indicated that a very satisfactory reproducibility was obtained.

Interference study

The anodic stripping analysis of Pb^{2+} could be affected by the interferences from other heavy metal ions. In this work, three metal ions including Cd^{2+} , Cu^{2+} and Hg^{2+} are selected as interfering ions to investigate the selectivity of the proposed method. Indeed, during the preconcentration of Cd^{2+} and Cu^{2+} these metal ions can occupy available deposition and/or coordination sites on the electrode surface which can decrease the peak current of Pb^{2+} . In the case of Hg^{2+} , the stripping peak current of Pb^{2+} can be affected by the formation of mercury film at the electrode surface, which can cause an easy reduction of Pb^{2+} and leading to a high peak current of Pb^{2+} due to the formation of an amalgam of Pb-Hg at the electrode surface. Fig.

9 shows the SWVs obtained using PPy/CNFs/CPE for the analysis of $50 \mu\text{g L}^{-1} \text{Pb}^{2+}$ in the absence and presence of different concentration metals ions interferents. It can be seen that the peak current of Pb^{2+} is almost constant in the presence of metals ions interferents, except in the case of the presence of Hg^{2+} a slight shift of peak potential of Pb^{2+} is observed. This marginal change of potential might be attributed to the formation of intermetallic compounds of Pb-Hg as reported in previous study [15, 60]. The obtained results clearly demonstrate that the selective detection of Pb^{2+} using PPy/CNFs/CPE under the optimized experimental conditions is not affected by the presence of metal ions interferents Cd^{2+} , Cu^{2+} , and Hg^{2+} .

Fig. 9 SWVs of 50 ppb Pb^{2+} in 0.1 M acetate buffer at PPy/CNFs/CPE in the absence (green line) and presence of interferents: $100 \mu\text{g L}^{-1}$ of Cd^{2+} (red line), and $50 \mu\text{g L}^{-1} \text{Cd}^{2+}$, $150 \mu\text{g L}^{-1}$ of Cu^{2+} and Hg^{2+} for each one (blue line)

Analytical application

The effectiveness application of this method to analysis Pb^{2+} ions in real samples was also tested by analyzing a tap water of our laboratory. For that, three samples of tap water were prepared. Briefly, an appropriate amount of acetic acid and sodium acetate were added to the water samples to obtain a concentration of 0.1 M for each compound, the pH was adjusted at 4.5. Then, the prepared samples were analyzed using the modified electrode PPy/CNFs/CPE under the optimized experimental conditions. The results showed the absence of obvious responses to Pb^{2+} , indicating that the concentration of Pb^{2+} in the analyzed samples was extremely low. In this case, the standard addition method was used, the three prepared samples were spiked with different concentrations of Pb^{2+} for recovery evaluation. As can be seen in Table 2, the recoveries are varied from 97.6% to 102%, indicating that the proposed method can be efficiently applied to Pb^{2+} detection in water real samples.

Table 2 Recovery and relative standard deviation (RSD) of Pb²⁺ in the tap water samples using the modified electrode PPy/CNFs/CPE (n = 3) (*Not detected)

Conclusion

In this work, a new nanocomposite of CNMs/CPs modified CPE was fabricated by combining the unique properties of polypyrrole and carbon nanofibers. The obtained nanocomposite provided excellent electrical conductivity and a high active surface area compared to the bare CPE as confirmed by CV and EIS measurements. Using SWASV and under the optimized experimental conditions, the modified electrode PPy/CNFs/CPE showed good analytical performances towards Pb^{2+} determination, with a very low detection limit, a satisfactory selectivity, an acceptable repeatability and reproducibility. Moreover, the proposed method was applied to the determination of Pb^{2+} in real water sample with satisfactory results, and also offers the possibility of simultaneously detection of several heavy metals ions Pb^{2+} , Cu^{2+} , Cd^{2+} , and Hg^{2+} .

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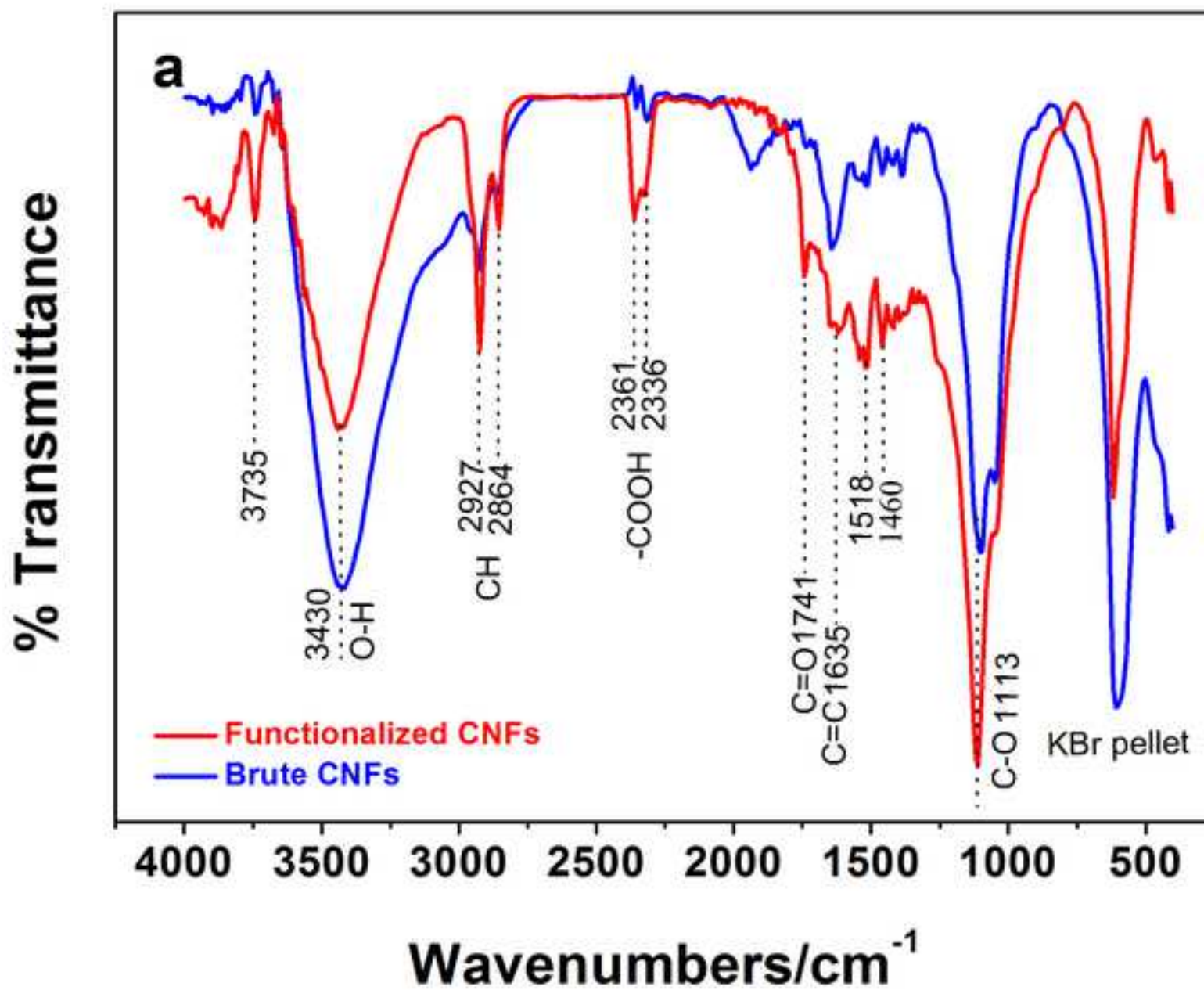
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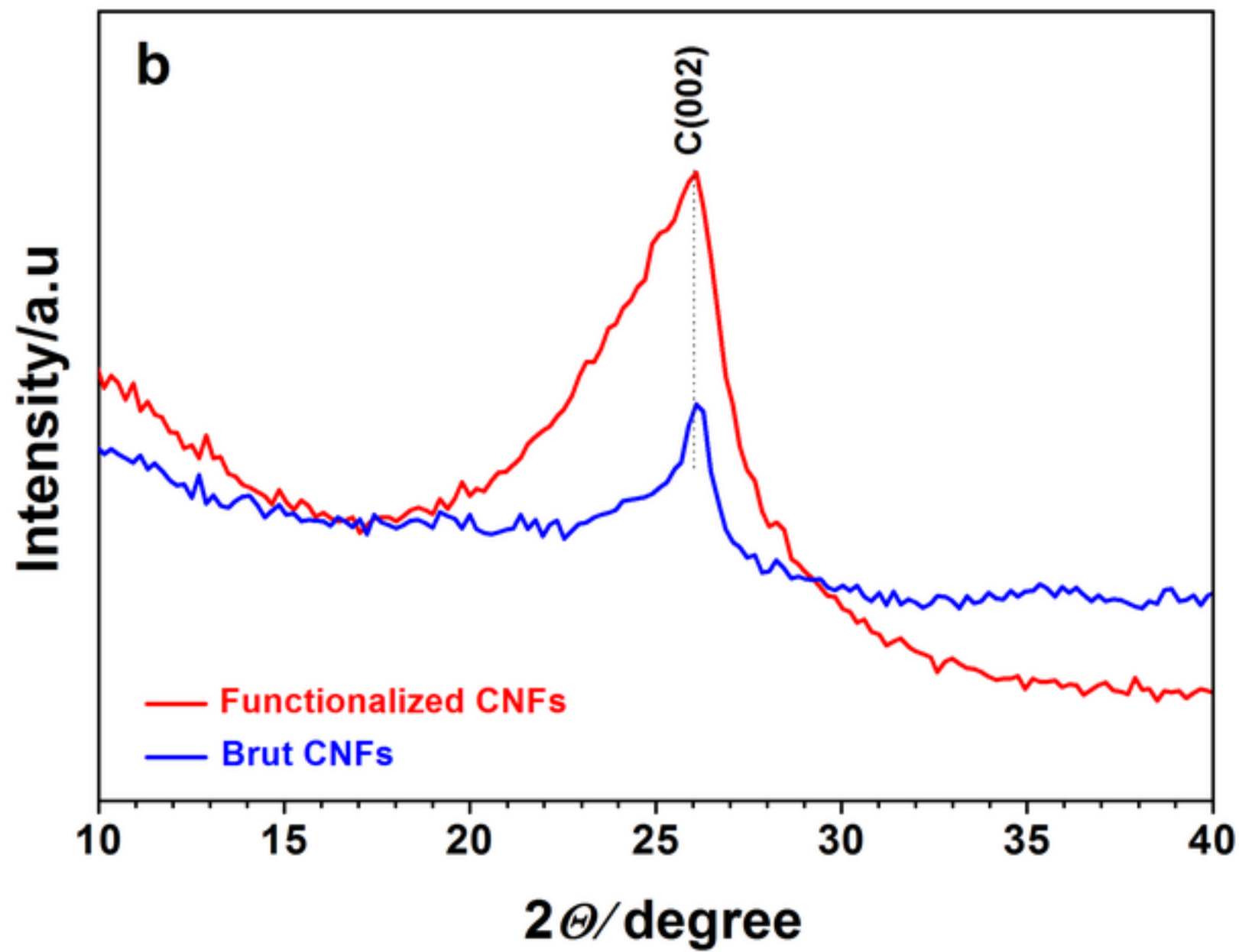
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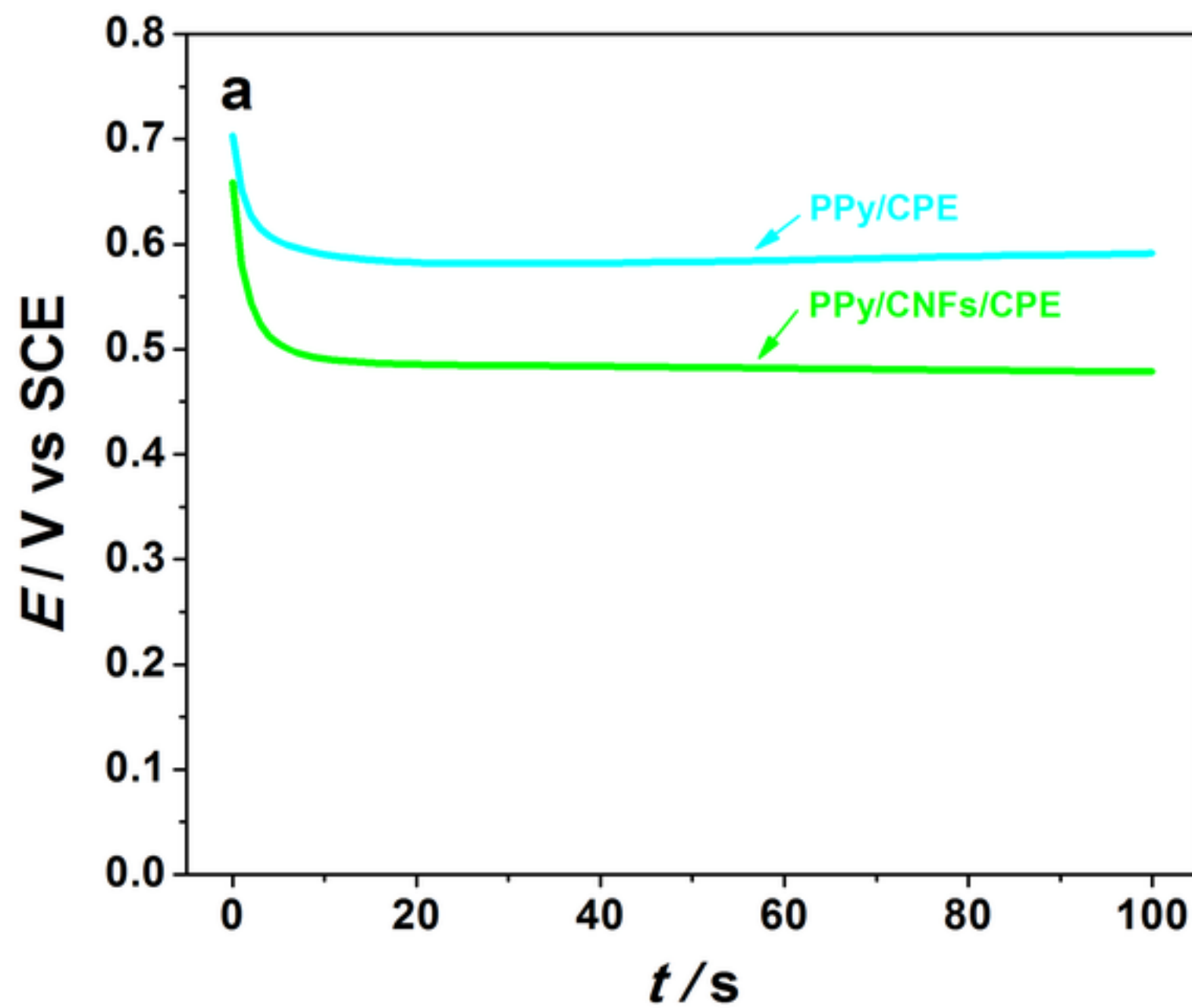
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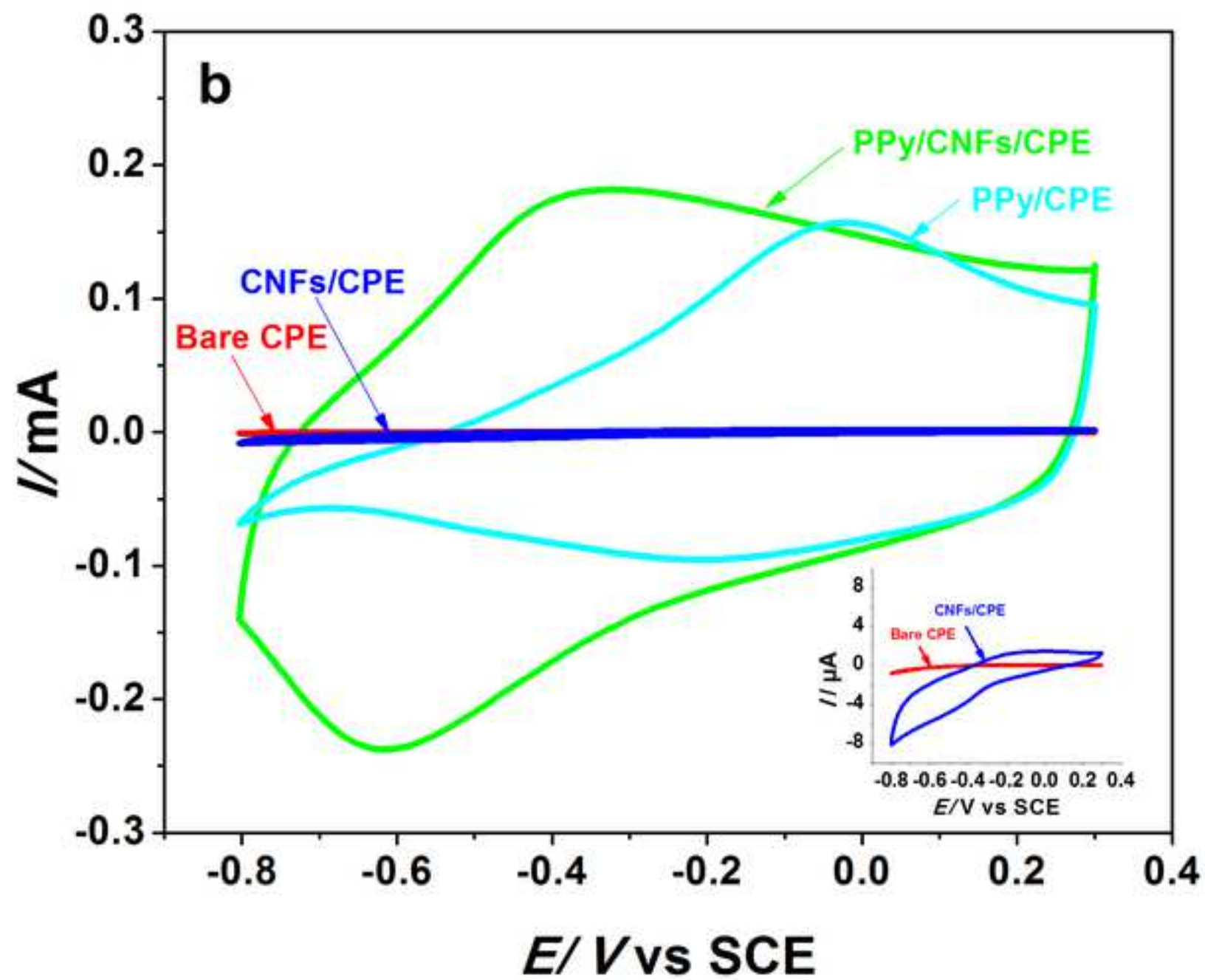
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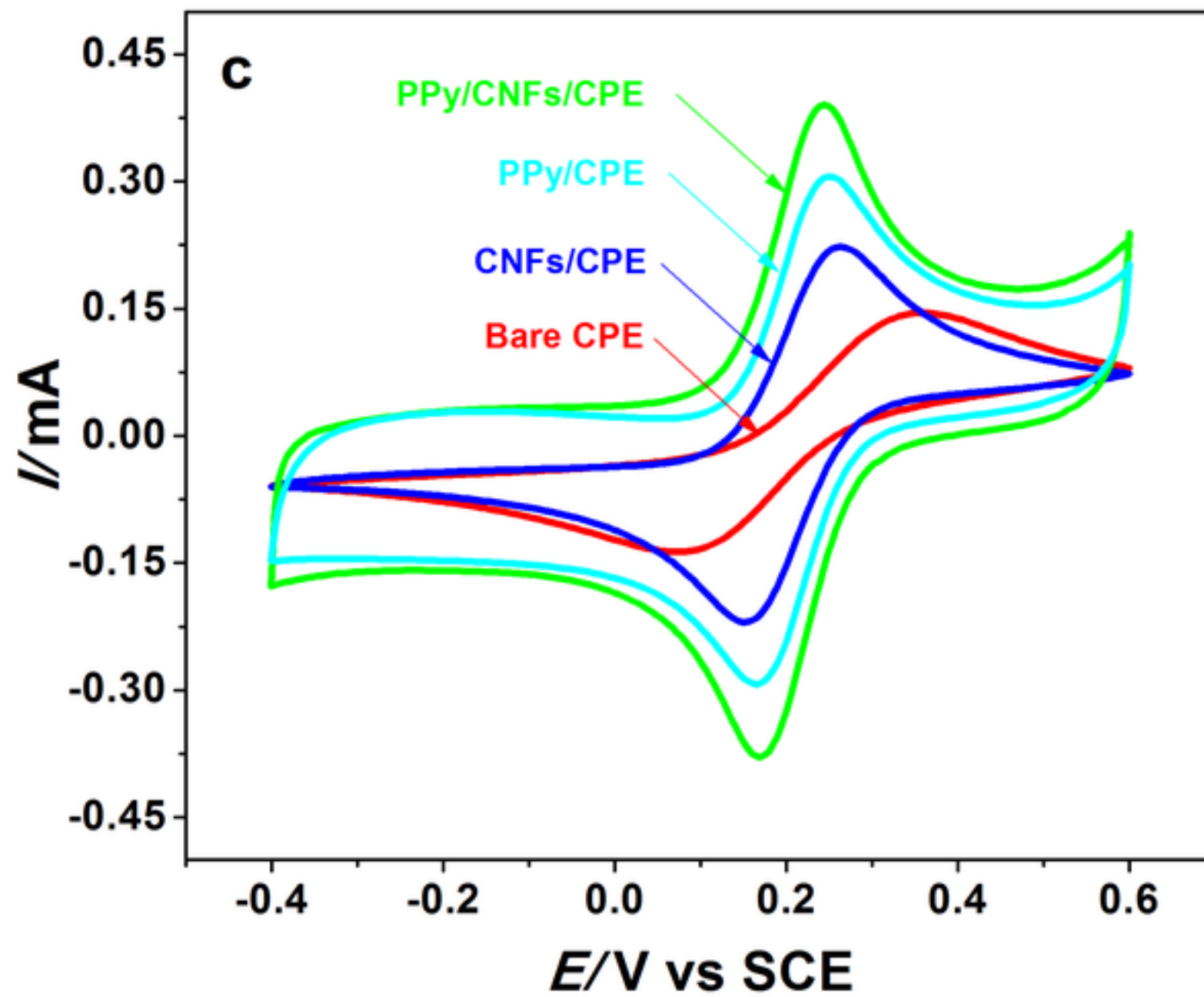


Fig. 3

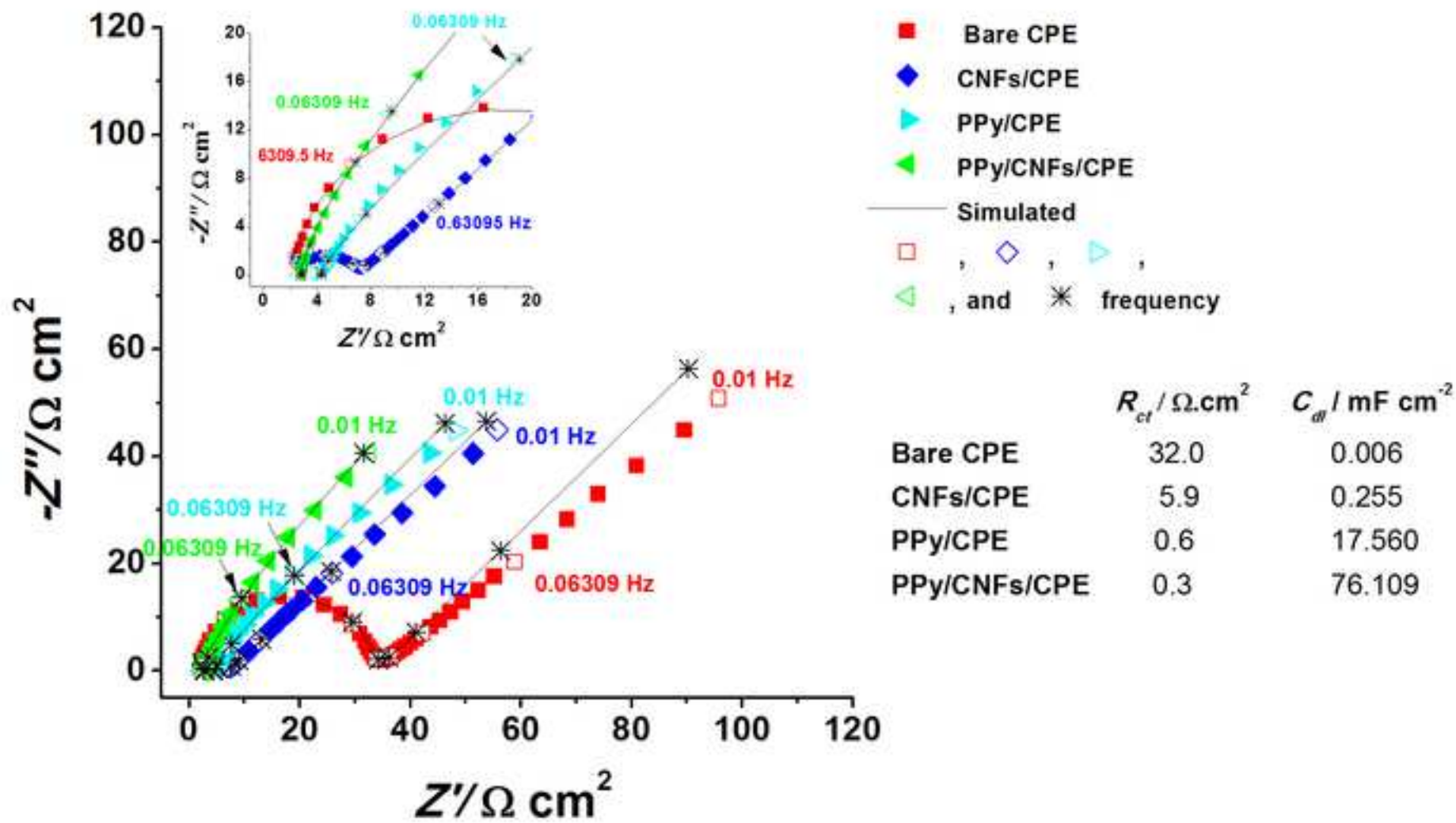
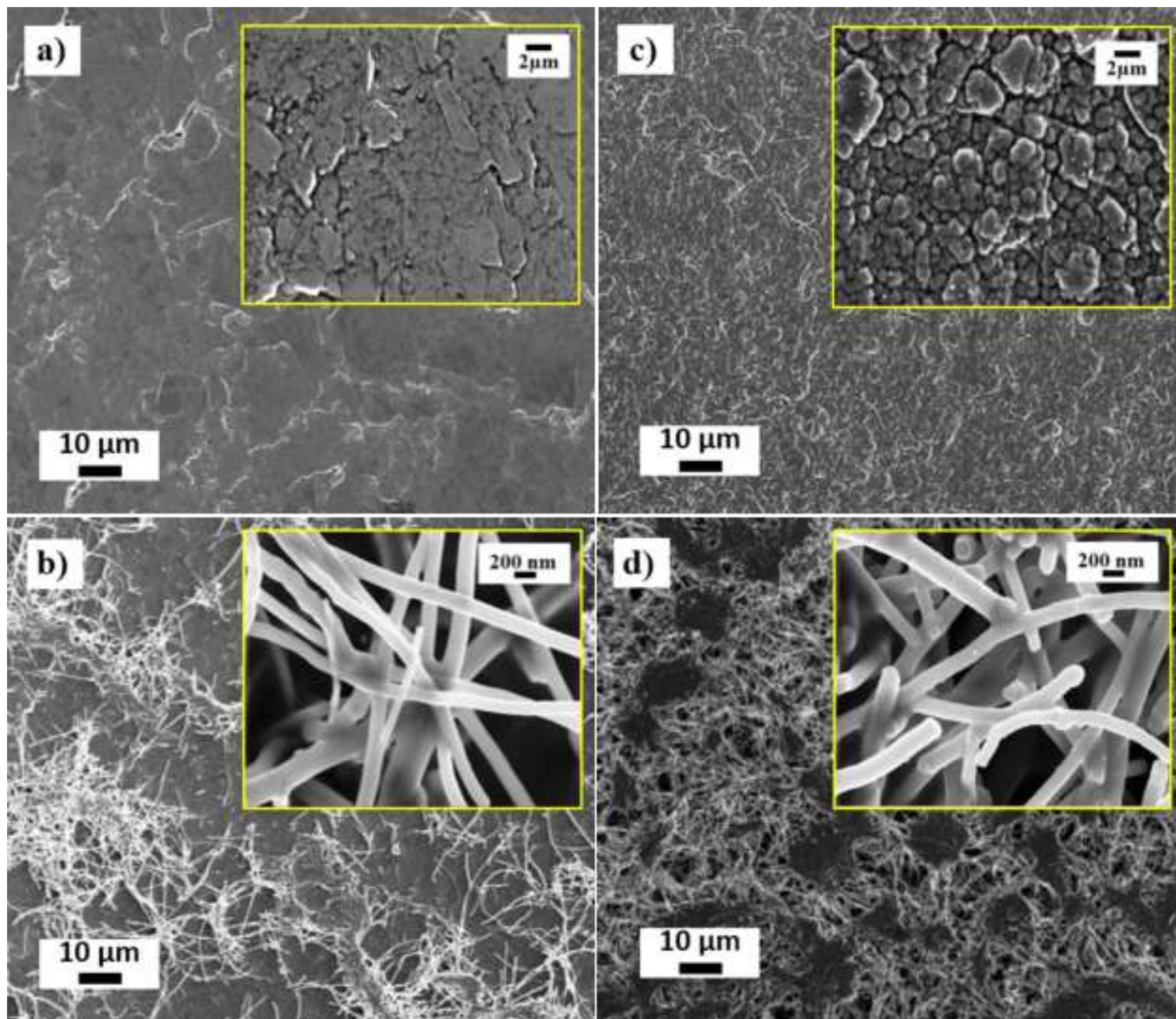


Fig. 4



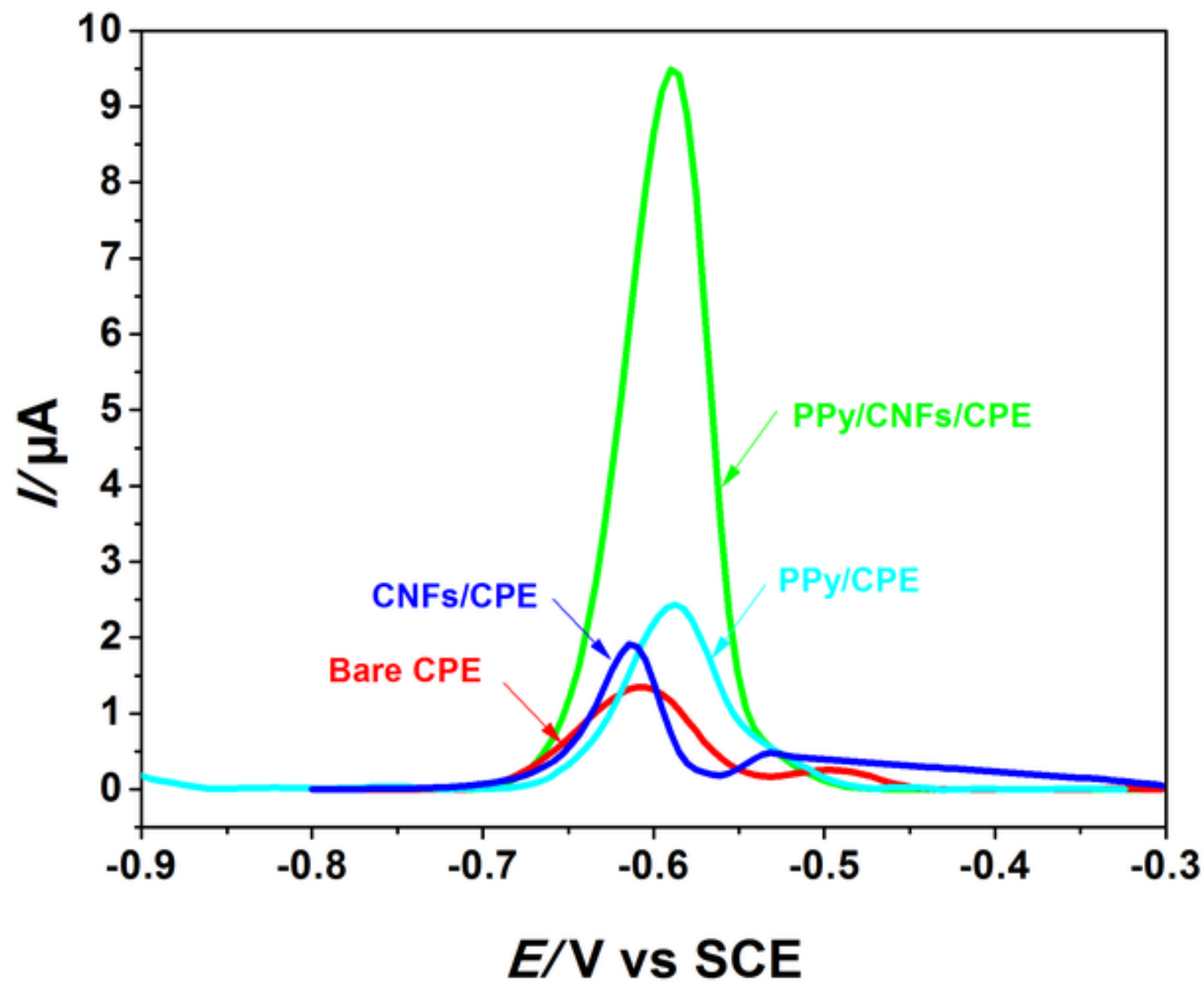


Fig. 6

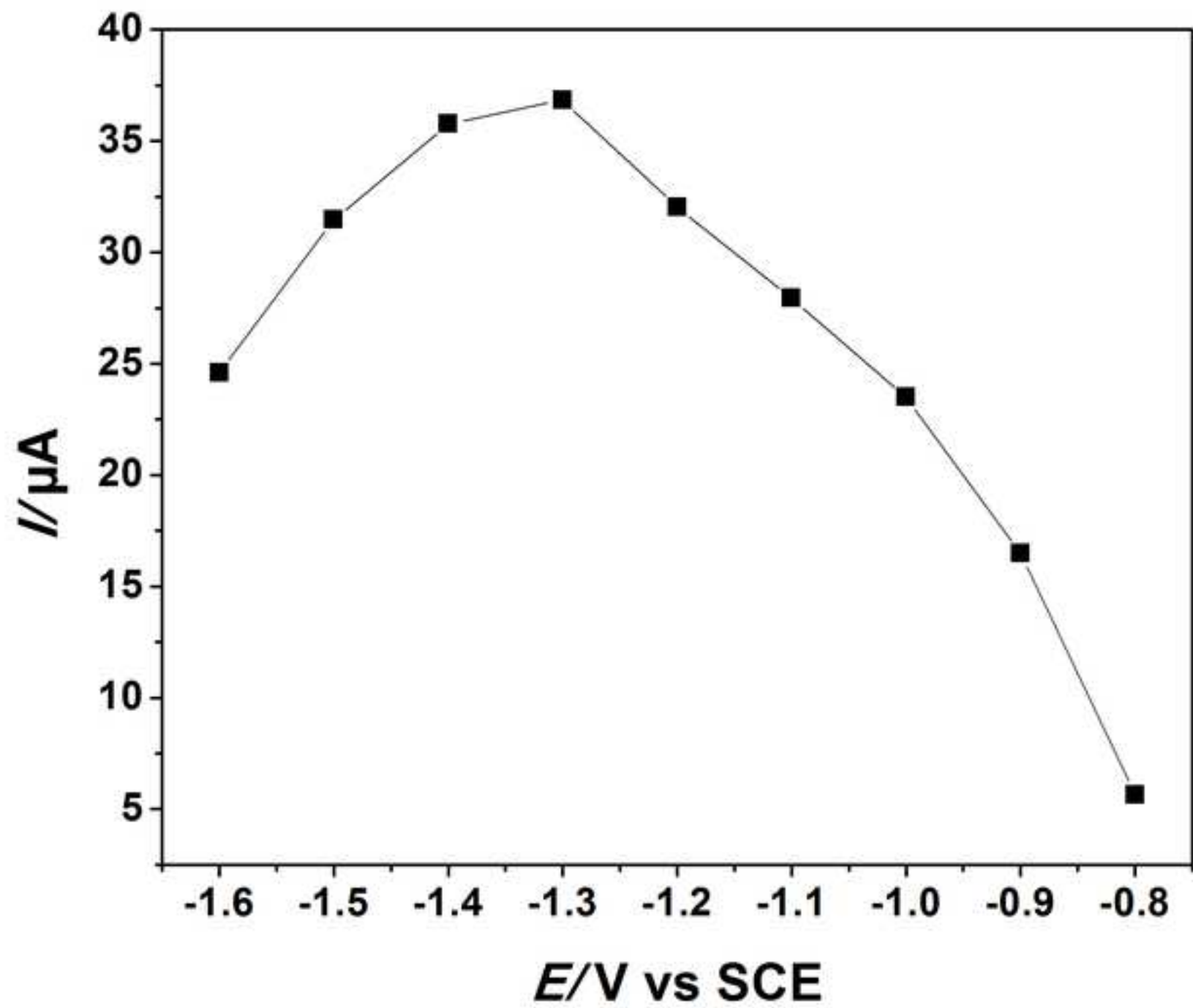


Fig. 7

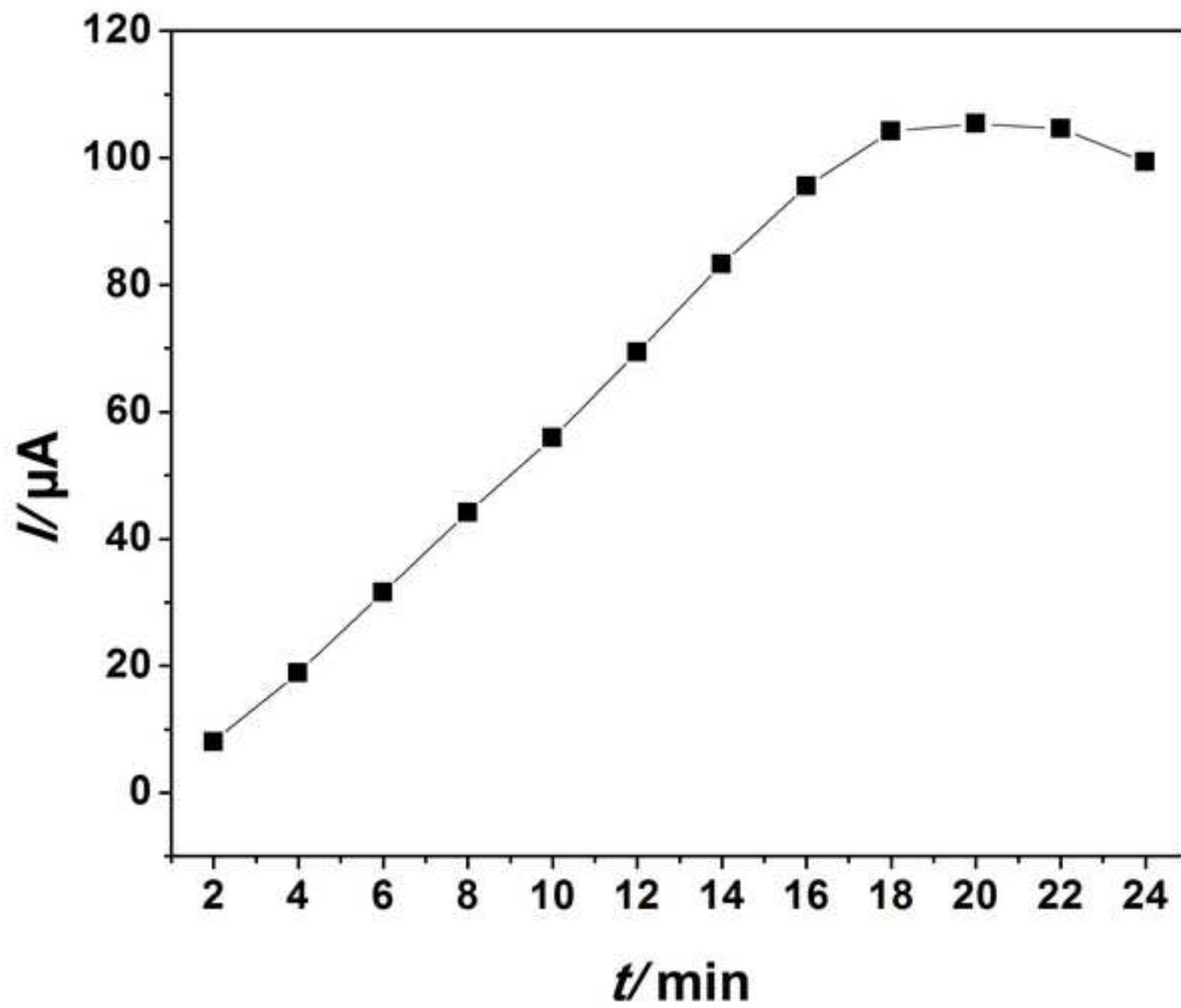
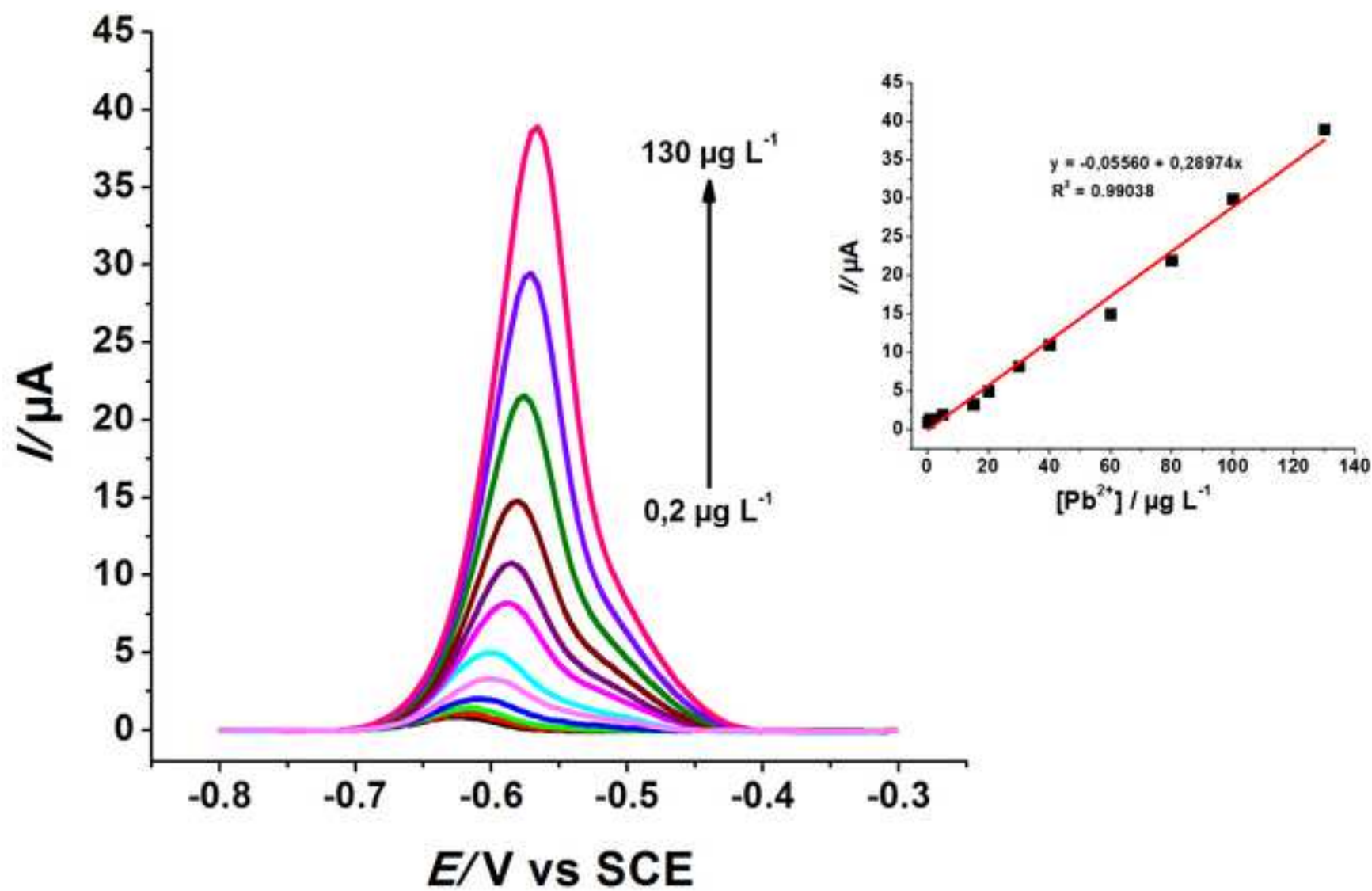


Fig. 8



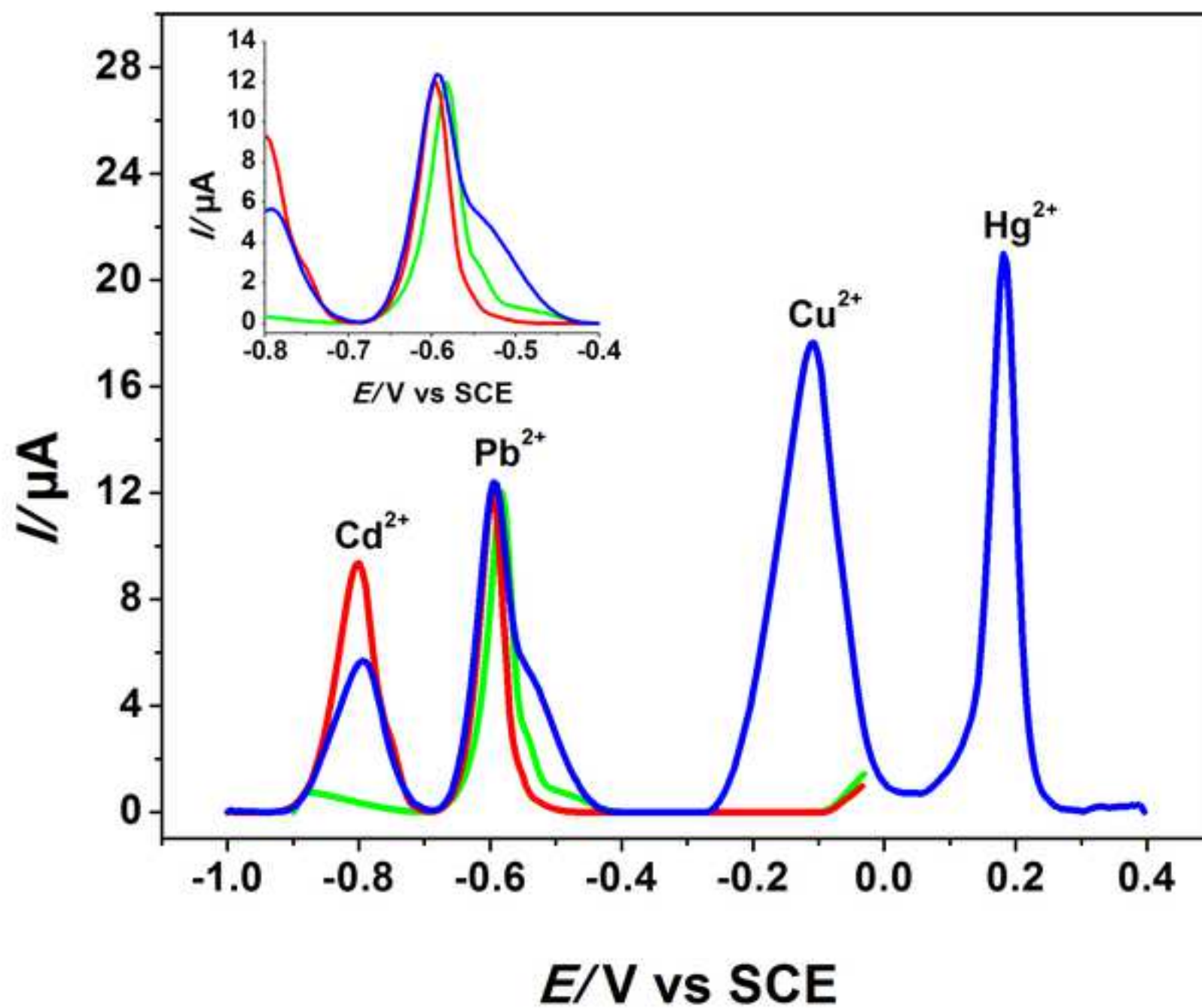


Table 1 Comparison of the analytical performance of PPy/CNFs/CPE with some other modified electrode for Pb²⁺ determination in aqueous media (detection limit and linear range are given in µg L⁻¹ except * in nM and \$ in µM)

Sensor	Technique	Deposition potential (V)	Deposition time (s)	Detection limit (µg L ⁻¹)	Linear range (µg L ⁻¹)	Ref
Poly(1,5-DAN)/MWCNTs/Pt	SWASV	-1.2	420	2.1	4-150	[11]
Bi/Nafion/OPPy-MES/GCE	SWASV	-1.2	300	0.03	0.05-35	[57]
AuNP-graphene-CySH-Bi-GCE	SWASV	-1.2	800	0.05	0.5-40	[58]
Poly(1,8-DAN)/CPE	DPASV	-0.9	600	30	40-2070	[53]
Bi-Poly(1,8-DAN)/CPE	SWASV	-1.0	120	0.3	0.5-50	[38]
Bi/PANI/GCE	SWASV	-1.3	120	16.5*	25-150*	[59]
CNFs/Nafion/ GCE	OSWSV	-1.25	300	0.9*	0.1-0.7\$	[51]
sGO/PPy-SPE	DPASV	-1.2	600	0.07	1.4-28, 28-280 and 280-14000	[15]
PPy/CNFs/CPE	SWASV	-1.2	600	0.05	0.2-130	This work

DPASV differential pulse anodic stripping voltammetry, *OSWSV* osteryoung square wave stripping voltammetry, *GCE* glassy carbon electrode, *SPE* screen printed electrode, *Pt* platinum, *1,5-DAN* 1,5-diaminonaphthalene, *1,8-DAN* 1,8-diaminonaphthalene, *OPPy* overoxidation of polypyrrole, *PANi* polyaniline, *MES* 2-mercaptoethanesulfonate, *sGO* cysteine functionalized graphene oxide, *MWCNTs* multiwalled carbon nanotubes, *CySH* L-cysteine, *AuNP* gold nanoparticles, *Bi* Bismuth

Table 2 Recovery and relative standard deviation (RSD) of Pb²⁺ in the tap water samples using the modified electrode PPy/CNFs/CPE (n = 3) (* Not detected)

Samples	Pb ²⁺ (µg L ⁻¹)			Recovery (%)	RSD (%)
	Original	Added	Found		
Tap water 1	ND*	5	4.88	97.6	1.2
Tap water 2	ND*	10	10.15	101.5	2.1
Tap water 3	ND*	15	15.30	102	2.3